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This chapter discusses that the pyrite oxidation in the natural environment is a surface controlled reaction and for this reason pyrite oxidation under such conditions may never come to completion. Pyrite oxidation models can be grouped into two categories depending on

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approach employed to make predictions:
deterministic functional models, and
mechanistic models.

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Mechanisms Microbial Role

Sulfide oxidation, part of sulfur's biotic/abiotic cycle, is an important natural phenomenon. However, because of the sulfide's association with metallic ores and fossil fuels in the form of pyrite (FeS_2) and the world's increasing demand

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for metals and fossil fuels, sulfide oxidation in nature is in some state of perturbation.

This perturbation, which results from land disturbances (e.g., mining, and/or ore processing), produces acid drainage often enriched with heavy metals.

A review: Pyrite oxidation mechanisms

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referred to as acid mine drainage and has

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become an economic and environmental burden this review deals with abiotic biotic modes of pyrite oxidation and the mechanistic

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Pyrite particles buried in aggregates

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cannot be coated with iron phosphate and can produce acid once the aggregates disintegrate. Coating and leaching-oxidation tests were carried out using 0.5 g of the pyritic shale sample mixed with approximately 1 g of sand and using a miscible displacement technique.

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complexity of the pyrite oxidation process.

Numerous elementary pathways are involved in the mechanism. The sulfur atoms have oxidation number -1 , and during the reactions, they end up with

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oxidation number +6, as sulfates.

Therefore, it is expected that several electron-transfer pathways might be accessible in the oxidative process. Basolo and

Pyrite Oxidation Mechanism by Oxygen
in Aqueous Medium

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It is now clear that the cathodic reaction is the rate-determining step for sulfide mineral oxidation (e.g., Brown and Jurinak, 1989). Williamson and Rimstidt (1994) showed that the pyrite oxidation rate depends on the concentration of Fe^{3+} or O_2 (Fig. 2, Fig. 3). Because the oxidant is involved in the activated

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complex, we can deduce that electron transfer from the sulfide to the oxidant, at

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Acid mine drainage (AMD) produced by
the oxidation of sulfide ore, especially

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pyrite, pollutes water and soil. Therefore, it is particularly important to control the oxidation of pyrite and the production of acid mine wastewater at the source.

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